


































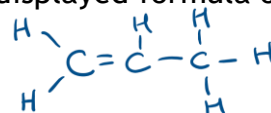
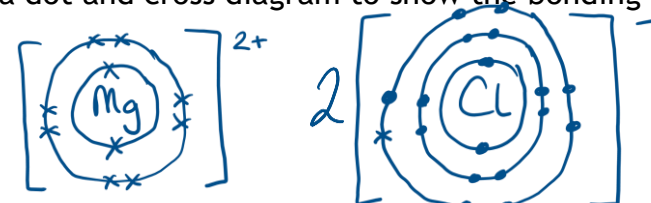
Retrieval Practice: Year 12 Number 2

Rules: Never look at your notes for retrieval practice! Do as many as you can, even if they are educated guesses. When you have tried (hard!) to answer them all, check the mark scheme and rate each question:

-  Easy, remembered perfectly
 Harder - could remember part of it or was familiar when I saw the answer
 Very hard - didn't recognise the answer so need to go back over this

	Question	Rating
1	Calculate the number of moles of Na_2CO_3 in 9.01 kg	  
2	Draw a displayed formula of propene	  
3	Give the formula of iron (III) nitrate	  
4	Write a balanced equation for the reaction between copper (II) oxide and hydrochloric acid	  
5	Draw a dot and cross diagram to show the bonding in magnesium chloride	  
6	Calculate the empirical formula of a compound containing 82.76% carbon and 17.24% hydrogen by mass	  
7	Describe the test for carbon dioxide gas	  
8	Balance the following equation: $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$	  
9	Calculate the M_r for $\text{Cr}_2(\text{SO}_4)_3$	  
10	Define oxidation and reduction in terms of electron transfer	  

Answers:

	Question															
1	Calculate the number of moles of Na_2CO_3 in 9.01 kg $M_r = 106$ $Mol = 9010 / 106 = 85$															
2	Draw a displayed formula of propene 															
3	Give the formula of iron (III) nitrate $\text{Fe}(\text{NO}_3)_3$															
4	Write a balanced equation for the reaction between copper (II) oxide and hydrochloric acid $\text{CuO} + 2\text{HCl} \rightarrow \text{CuCl}_2 + \text{H}_2\text{O}$															
5	Draw a dot and cross diagram to show the bonding in magnesium chloride 															
6	Calculate the empirical formula of a compound containing 82.76% carbon and 17.24% hydrogen by mass <table><tr><td>C</td><td>H</td><td></td></tr><tr><td>$82.76 / 12$</td><td>$17.24 / 1$</td><td></td></tr><tr><td>6.897</td><td>17.24</td><td></td></tr><tr><td>1</td><td>2.5</td><td></td></tr><tr><td>2</td><td>5</td><td>C_2H_5</td></tr></table>	C	H		$82.76 / 12$	$17.24 / 1$		6.897	17.24		1	2.5		2	5	C_2H_5
C	H															
$82.76 / 12$	$17.24 / 1$															
6.897	17.24															
1	2.5															
2	5	C_2H_5														
7	Describe the test for carbon dioxide gas (Bubble through) limewater - turns cloudy/milky															
8	Balance the following equation: $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$ $2\text{FeS}_2 + 5.5\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + 4\text{SO}_2$ OR $4\text{FeS}_2 + 11\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 8\text{SO}_2$															
9	Calculate the M_r for $\text{Cr}_2(\text{SO}_4)_3$ 392.3															
10	Define oxidation and reduction in terms of electron transfer Oxidation = loss of electrons, reduction = gain of electrons															